

DATE:

NAME:

CLASS:

BLM 2-43
continued

2. Complete the following table by filling in the missing information about isotopes. The first row is completed as an example.

Name of Isotope	Symbol	Mass Number	Number of Protons	Number of Neutrons
hydrogen-3	${}^3_1\text{H}$	3	1	2
scandium-49				
	${}^{60}_{27}\text{Co}$			
nitrogen-15				
	${}^{238}_{92}\text{U}$			
	${}^{129}_{53}\text{I}$			
			56	79
		86		48
	${}^{18}_8\text{O}$			
carbon-14				

3. Although oxygen-16 is the most common isotope of oxygen, oxygen-17 and oxygen-18 are also present. Despite the differences in the atomic structures of the three isotopes, there is no difference in how they form ionic or covalent compounds with atoms of other elements. Explain how this can be.
