

Use with textbook pages 186–196.

Chemical names and formulas of ionic compounds

1. Write the name for each of the following compounds.

- (a) BeS beryllium sulphide
(b) Hg_3N_2 mercury (II) nitride
(c) $\text{Cu}(\text{NO}_3)_2$ copper (II) nitrate
(d) Ag_2O silver oxide
(e) CoBr_2 cobalt (II) bromide
(f) $\text{Bi}_3(\text{PO}_4)_5$ bismuth (V) phosphate
(g) CaF_2 calcium fluoride
(h) Mn_2O_3 manganese (III) oxide
(i) $\text{Cr}_2(\text{SO}_4)_3$ chromium (III) sulphate
(j) ZnCl_2 zinc chloride

- (k) $\text{Ni}(\text{OH})_2$ nickel (II) hydroxide
(l) $\text{K}_2\text{Cr}_2\text{O}_7$ potassium dichromate
(m) ScF_3 scandium fluoride
(n) NaI sodium iodide
(o) $\text{Pb}(\text{CO}_3)_2$ lead (IV) carbonate
(p) RbClO_2 rubidium chlorite
(q) K_3P phosphorus potassium phosphate
(r) $\text{Mg}(\text{CN})_2$ magnesium cyanide
(s) SnS tin (II) sulphide
(t) NaHCO_3 sodium bicarbonate

2. Write the chemical formula for each of the following compounds.

- (a) aluminum bromide AlBr_3
(b) platinum(II) sulphide PtS
(c) strontium sulfite SrSO_3
(d) scandium oxide Sc_2O_3
(e) titanium(IV) nitride $\text{Ti}(\text{NO}_2)_4$
(f) ammonium sulphate $(\text{NH}_4)_2\text{SO}_4$
(g) lithium selenide Li_2Se
(h) lead(II) hydrogen sulphate $\text{Pb}(\text{HSO}_4)_2$
(i) sodium acetate NaCH_3COO
(j) cesium chloride CsCl

- (k) cadmium(II) hydroxide $\text{Cd}(\text{OH})_2$
(l) zinc phosphate $\text{Zn}_3(\text{PO}_4)_2$
(m) barium chloride BaCl_2
(n) tin(II) permanganate $\text{Sn}(\text{MnO}_4)_2$
(o) lithium hypochlorite LiClO
(p) gold(III) sulphate $\text{Au}_2(\text{SO}_4)_3$
(q) sodium nitrate NaNO_3
(r) chromium(III) chloride CrCl_3
(s) potassium carbonate K_2CO_3
(t) iron(III) bisulphate $\text{Fe}(\text{HSO}_4)_3$