

Balance the reactions and indicate which types of chemical reaction are being represented:

- a)  $\text{NaBr} + \text{Ca(OH)}_2 \rightarrow \text{CaBr}_2 + \text{NaOH}$  Reaction Type : \_\_\_\_\_
- b)  $\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$  Reaction Type : \_\_\_\_\_
- c)  $\text{C}_5\text{H}_9\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  Reaction Type : \_\_\_\_\_
- d)  $\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$  Reaction Type : \_\_\_\_\_
- e)  $\text{Li}_3\text{N} + \text{NH}_4\text{NO}_3 \rightarrow \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$  Reaction Type : \_\_\_\_\_
- f)  $\text{HBr} + \text{Al(OH)}_3 \rightarrow \text{H}_2\text{O} + \text{AlBr}_3$  Reaction Type : \_\_\_\_\_
- g)  $\text{Na}_3\text{PO}_4 + \text{KOH} \rightarrow \text{NaOH} + \text{K}_3\text{PO}_4$  Reaction Type \_\_\_\_\_
- h)  $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + \text{LiCl}$  Reaction Type \_\_\_\_\_
- i)  $\text{C}_6\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  Reaction Type \_\_\_\_\_
- j)  $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$  Reaction Type \_\_\_\_\_
- k)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  Reaction Type \_\_\_\_\_
- l)  $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_3$  Reaction Type \_\_\_\_\_
- m)  $\text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + \text{RbF}$  Reaction Type \_\_\_\_\_
- n)  $\text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + \text{Ag}$  Reaction Type \_\_\_\_\_
- o)  $\text{C}_3\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  Reaction Type \_\_\_\_\_
- p)  $\text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe(C}_5\text{H}_5)_2$  Reaction Type \_\_\_\_\_
- q)  $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + \text{Cl}_2$  Reaction Type \_\_\_\_\_
- r)  $\text{MgI}_2 + \text{Mn(SO}_3)_2 \rightarrow \text{MgSO}_3 + \text{MnI}_4$  Reaction Type \_\_\_\_\_
- s)  $\text{O}_3 \rightarrow \text{O} \cdot + \text{O}_2$  Reaction Type \_\_\_\_\_
- t)  $\text{NO}_2 \rightarrow \text{O}_2 + \text{N}_2$  Reaction Type \_\_\_\_\_