Name: _____ Date: _____ Period: 1 2 3 4 5 6 7

**Types of Chemical Reaction and Predicting Products Worksheet**

*Balance the following. Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement, acid-base, combustion, precipitation, or gas evolution) is being represented in 1 to 13.*

1. 1 $Na_3PO_4$ (aq)+ 3 KOH(aq) → 3 NaOH(aq) + 1 $K_3PO_4$(s)
   Reaction Type DR
2. 1 $MgCl_2$ (S) + 1 $Li_2CO_3$(s) → 1 $MgCO_3$(s) + 2 LiCl(aq)
   Reaction Type DR
3. 1 $C_6H_{12}$(l) + 9 $O_2$(g) → 6 $CO_2$(g) + 6 $H_2O$(l)
   Reaction Type C
4. 1 Pb(s) + 1 $FeSO_4$(aq) → 1 $PbSO_4$(s) + 1 Fe(s)
   Reaction Type SR
5. 1 $CaCO_3$(s) → 1 CaO(s) + 1 $CO_2$ (g)
   Reaction Type D
6. 1 $P_4$(s) + 3 $O_2$(g) → 2 $P_2O_3$(s)
   Reaction Type S
7. 2 $RbNO_3$(aq) + 1 $BeF_2$(aq) → 1 $Be(NO_3)_2$(aq) + 2 RbF(aq)
   Reaction Type DR
8. 2 $AgNO_3$(aq) + 1 Cu(s) → 1 $Cu(NO_3)_2$(aq) + 2 Ag(s)
   Reaction Type SR
9. 1 $C_3H_6O$(l) + 4 $O_2$(g) → 3 $CO_2$(g) + 3 $H_2O$(l)
   Reaction Type C
10. 2 $C_5H_5$(l) + 1 Fe(s) → 1 $Fe(C_5H_5)_2$ (s)
    Reaction Type S
11. 1 $SeCl_6$ (s)+ 1 $O_2$(g) → 1 $SeO_2$(s) + 3 $Cl_2$(g)
    Reaction Type SR
12. 2 $MgI_2$(aq) + 1 $Mn(SO_3)_2$(s) → 2 $MgSO_3$(s) + 1 $MnI_4$(aq)
    Reaction Type DR
13. 2 $NO_2$(g) → 2 $O_2$(g) + 1 $N_2$(g)
    Reaction Type S